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(54) Novel faujasite-type aluminosilicates, a method of producing the same, and hydrocracking catalysts for heavy hydrocarbon oils.

(57) A hydrocracking catalyst for heavy hydrocarbon oils comprising a metallic element of the Vlb group and a metallic element of the VIII Group supported on a carrier containing a novel faujasite-type aluminosilicate which absorbs an infrared in a frequency region of $3740 \pm 10 \text{ cm}^{-1}$ in an absorption percentage A of at least 20 % and absorbs an infrared in a frequency region of $3560 \pm 10 \text{ cm}^{-1}$ in an absorption percentage B of at least 5 %, the ratio of A/B being at least 2, has a specific surface area of at least $650 \text{ m}^2/\text{g}$, has a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 20 to 50, and has a lattice constant of from 24.15 to 24.50 \AA . The novel faujasite-type aluminosilicate is produced by treating a faujasite-type zeolite with an acid.

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and as well excel in the yields of light hydrocarbon oils such as gas oil, kerosine, and naphtha, and they have repeated researches as to the properties and producing method of the zeolites to be used as carriers or materials suitable for such catalysts.

It is known, as described above, that the catalytic properties of the conventional hydrocracking catalysts of the kinds described above vary considerably depending on the parent zeolites (steaming zeolites) used as carriers and the conditions of the acid treatment. Optimization of the catalytic properties including cracking and hydrogenation efficiency is necessary for performing the hydrocracking of heavy hydrocarbon oils efficiently with a sufficient yield of light hydrocarbon oils such as gas oil, kerosine, and naphtha, and such an optimization requires further development of carriers having properties enabling the optimization.

The inventors directed their attention mainly to the point described above, and they prepared zeolites having various properties by conducting the acid treatment using parent zeolites of various types and properties under various conditions and, then, investigated in detail how the properties of the zeolites relate to the heavy hydrocarbon oil hydrocracking performance.

As the result, they found that metallic ingredients having hydrogenation activity can be supported with good dispersibility on novel zeolites (faujasite-type zeolites) which are characterized by that the absorption percentages at specified frequency region of infrared absorption spectrum, the framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio, and the specific surface area are in specified ranges. respectively, in spite of the high acidity of the novel zeolites. They also found that hydrocracking catalysts satisfying the above objects of the present invention such as catalytic activity are obtainable by supporting specified hydrocracking metallic ingredients on the novel zeolites, and eventually, they completed the present invention.

That is, the present invention provides a novel faujasite-type aluminosilicate which absorbs an infrared in a frequency region of $3740 \pm 10 \text{ cm}^{-1}$ in an absorption percentage A of at least 20 and absorbs an infrared in a frequency region of $3560 \pm 10 \text{ cm}^{-1}$ in an absorption percentage B of at least 5 %, the ratio of A/B being at least 2, has a specific surface area of at least $650 \text{ m}^2/\text{g}$, has a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 20 to 50, and has a lattice constant of from 24.15 to 23.50 \AA .

The present invention further provides a suitable method of producing the faujasite-type aluminosilicate catalyst described above, which comprises treating a faujasite-type zeolite with from 2 to 20 moles of an acid per 1 kg of said faujasite-type zeolite, said faujasite-type zeolite having a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 15 to 25, having a specific surface area of at least $500 \text{ m}^2/\text{g}$, and having a lattice constant of from 24.36 to 24.50 \AA .

Further, the present invention relates to the application of the faujasite-type aluminosilicate and provides a hydrocracking catalyst for a heavy hydrocarbon oil comprising a metallic element belonging to the VIb Group of the Periodic Table and a metallic element belonging to the VIII Group of the Periodic Table which is supported on a carrier containing the faujasite-type aluminosilicate.

THE PREFERRED EMBODIMENTS OF THE INVENTION

The percentage of infrared absorption was determined by measuring the infrared absorption spectrum of a film (disc) of about 0.1 mm in thickness made of from 20 to 30 mg of a sample using Self-Support method and then calculating the percentage of infrared absorption by using the high frequency side end of the absorption region at 3740 cm^{-1} as the standard (base line).

The framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio was determined by measuring the areas of peaks in high resolution solid Si-NMR spectrum chart in accordance with the general method and calculating the molar ratio using the areas and the following equation:

Skeletal $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio

$$= 2 \times \sum_{n=0}^4 \text{Isi}(n\text{Al}) / \sum_{n=0}^4 (n/4) \text{Isi}(n\text{Al})$$

wherein $\text{Isi}(n\text{Al})$ represents the area of the peak of $\text{Si}(n\text{Al})$, n representing the number of Al atoms bonding to a Si atom through oxygen atoms (coordination number) and representing an integer of 0 to 4, and the NMR chemical shifts of $\text{Si}(\text{OAl})$, $\text{Si}(1\text{Al})$, $\text{Si}(2\text{Al})$, $\text{Si}(3\text{Al})$, and $\text{Si}(4\text{Al})$ corresponding respectively to the coordination numbers are $-86 \pm 3 \text{ ppm}$, $-80 \pm 3 \text{ ppm}$, $-75 \pm 3 \text{ ppm}$, $-67 \pm 3 \text{ ppm}$, and $-61 \pm 3 \text{ ppm}$, respectively, when the internal standard is Si in a silicon rubber. This measuring method enables to determine the

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$\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio only in the zeolite framework, while the ordinary elementary analysis determines a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio which further includes the content of the aluminum being away from the zeolite framework.

The specific surface area was determined by using a BET method by means of N_2 adsorption in accordance with the ordinary method.

The lattice constant was determined from diffraction angle with X-ray diffraction.

Although the method to be used for producing the faujasite-type aluminosilicate of the present invention is not particularly limited, the method of the present invention is usually suitable.

Hereinafter, the method of the present invention will be described in detail.

The material to be used in the method of the present invention for producing the faujasite-type aluminosilicate is a faujasite-type zeolite having a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 15 to 25, preferably from 17 to 23, a specific surface area of at least $500 \text{ m}^2/\text{g}$, preferably at least $520 \text{ m}^2/\text{g}$, and a lattice constant of from 24.36 to 24.50 \AA , preferably from 24.37 to 24.45 \AA .

The preferred faujasite-type zeolite to be used as the starting material is a steaming faujasite-type zeolite (a faujasite-type zeolite which has been treated in hydrothermal condition), particularly, a steaming faujasite-type zeolite of high crystallinity. As regards zeolites of this type, it is known that the higher the crystallinity is, the larger the specific surface area is, and the larger the specific surface area is, the higher the crystallinity is.

Faujasite-type zeolites having a lattice constant of more than 24.50 \AA are poor in acid-resistance and will suffer from ready destruction of the crystallinity into undesirable form, and, on the other hand, faujasite-type zeolites having a lattice constant of less than 24.36 \AA will cause a decrease in the acidity itself to be used and make it difficult to obtain catalysts having sufficient hydrocracking activity.

The steaming faujasite-type zeolite described above is obtainable by using, as a material, various kinds of faujasite-type zeolites having a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of at least 4.8, preferably at least 5.0 and a Na_2O content decreased to 3.0 % by weight or less, preferably to 2.0 % by weight or less, by means of NH_4^+ ion exchange, and treating them with steam under the following conditions in accordance with the ordinary hydrothermal treatment.

That is, the conditions appropriate for the hydrothermal treatment generally include a treating temperature of from 500 to 900°C , preferably from 520 to 850°C , a treating time of from 0.5 to 5 hours, preferably from 1.0 to 4.5 hours, and a partial pressure of steam of from 1 to 100 %, preferably from 5 to 100 %. Also, self-steaming using the water possessed in the zeolite used may be employed. The steaming treatment may be conducted by use of various kinds of systems including flow system and closed system.

In the method of the present invention, among the above-described faujasite-type zeolites including the steaming faujasite-type zeolite obtained by the above steaming treatment, the faujasite-type zeolite having the above-described specified properties is treated with an acid.

The treatment with an acid is so conducted as to remove efficiently not only the alumina which is not a constituent of the zeolite framework, for example, the alumina layer disconnected from the zeolite framework by the steaming treatment, but also a portion of the aluminum (oxide) remaining as a constituent of the zeolite framework.

Although the aluminum (oxide) in the zeolite framework can be disconnected by the steaming treatment or other means, it is important to perform the disconnection and removal excessively by the treatment with the acid as described above.

That is, in order to provide the specified novel properties to the surface of the zeolite, it is just effective to disconnect and remove even a part of the aluminum (oxide) in the zeolite skeleton by means of the above-described treatment with the acid.

In concrete terms, the treatment with the acid as described above increases sufficiently the content of the silanol group (Si-OH) which causes the infrared absorption in a frequency region ($3740 \pm 10 \text{ cm}^{-1}$).

However, excessive acid treatment may induce undesirable destruction of zeolite crystals, resulting in a decrease in the specific surface area, and will cause a decrease in the amount of the silanol groups (Si-OH) and a decrease beyond the necessity in the amount of the aluminum which takes part in generating the hydrocracking activity.

The acid treatment, therefore, is conducted so moderately as to keep the specific surface area of the zeolite at least to $650 \text{ m}^2/\text{g}$, preferably at least to $660 \text{ m}^2/\text{g}$, the absorption percentage A in the region of infrared frequencies ($3740 \pm 10 \text{ cm}^{-1}$) due to the silanol group (Si-OH) at least to 20 %, preferably at least to 22 %, the absorption percentage B in the region of infrared frequencies ($3560 \pm 10 \text{ cm}^{-1}$) due to the active point (Brønsted acid) of the aluminum which takes part in generating the hydrocracking activity at least to 5 %, preferably at least to 7 %, and the ratio between these absorption percentages, A/B, at least to 2, preferably at least to 2.1.

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Metal elements belonging to each Group may be used individually or in a combination of two or more of them.

The desirable amount of the metal element belonging to the VIb Group is generally from 3 to 24 % by weight, preferably from 8 to 20 % by weight, based on the total weight of the catalyst, and the desirable amount of the metal element belonging to the VIII Group is generally from 0.7 to 20 % by weight, preferably from 1.5 to 8 % by weight, based on the total weight of the catalyst.

At the time of supporting the metallic ingredient belonging to the VIb Group and the VIII Group respectively on the carrier, supporting of each metallic element may be carried out by using a compound which is used for the preparation of conventional supported catalysts. Element metals or alloys also may be used.

Some examples of the compound of molybdenum or tungsten include molybdic acid, tungstic acid, molybdic acid salts such as ammonium molybdate, tungstic acid salts such as ammonium tungstate, heteropoly-acids containing molybdenum and/or tungsten, salts of the heteropoly-acids, and various complex compounds such as molybdenum carbonyl and tungsten carbonyl. Among these, the particularly preferred are ammonium molybdate and ammonium tungstate.

Some examples of the compound of cobalt or nickel include inorganic acid salts such as nitrates, sulfates, phosphates, carbonates, and chlorides, organic acid salts such as acetates, and various inorganic and organic complex compounds such as amine complexes and carbonyl complexes, among these, the particularly preferred are cobalt nitrate and nickel nitrate.

Other additives such as other metallic ingredients may be added, according to demand, to the extent that the accomplishment of the object of the present invention is not hindered.

The method for supporting is not particularly limited, and various methods may be used, for example, wet supporting techniques using proper solutions, such as an aqueous solution or slurries, including impregnation (including vacuum impregnation), coprecipitation, wet-kneading, adsorption, ion-exchange, and spraying, dry-supporting techniques, including mechanical mixing, vapor-phase adsorption, evaporation, and sublimation, and a combination thereof. Usually, the supporting may be carried out according to an ordinary technique, such as coprecipitation, impregnation or kneading.

After the supporting of the metallic ingredients described above, drying and calcination are carried out optionally in accordance with ordinary methods, followed by, if desired, activation or stabilization such as reduction treatment or pre-sulfurization treatment, to complete the production of the hydrocracking catalyst for heavy hydrocarbon oils of the present invention. These treatments of various kinds may also be performed in the reaction vessel for the hydrocracking previous to the reaction.

The hydrocracking catalyst for heavy hydrocarbon oils of the present invention is an excellent catalyst with a long life, high activity, and high selectivity for light hydrocarbon oils such as gas oil, kerosine, and naphtha and, therefore, may be suitably used in various fields of hydrocracking of heavy hydrocarbon oils, particularly in the field of producing light hydrocarbon oils such as gas oil, kerosine, and naphtha. Also, the catalysts may be used as a catalyst or a catalytic ingredient for other relating hydrogenation treatments.

The following description concerns the reaction conditions of the hydrocracking of heavy hydrocarbon oils under which the hydrocracking catalysts for heavy hydrocarbon oils of the present invention can work most efficiently (reaction conditions for hydrocracking heavy hydrocarbon oils so efficiently as to produce light hydrocarbon oils such as kerosine and naphtha in a high yield).

Some examples of the heavy hydrocarbon oil which is a material suitable for the hydrocracking reaction include atmospheric distillation residual oil, vacuum distillation residual oil, vacuum heavy gas oil, clarified oil, heavy cycle oil, visbreaking oil, tar sand oil, and shale oil.

These may be used individually or as a mixture of two or more of them.

The reaction conditions may be selected from wide ranges including the reaction conditions hitherto undertaken in the conventional hydrocracking of heavy hydrocarbon oils, and usually, the preferable reaction temperature is from 350 to 450 °C, the preferable reaction pressure is from 50 to 200 kg/cm², the preferable ratio between the hydrogen gas supplied and the material oil supplied [H₂/oil] is from 500 to 5,000 Nm³/kl, and the preferable LHSV is from 0.1 to 10 hr⁻¹.

The reaction system for the hydrocracking is not particularly limited and may be selected from various known systems including continuous flow system using fixed bed, moving bed, fluidized bed or suspensoid bed, semi-continuous system, and batch system. Usually, continuous flow system using, for example, fixed bed may be suitably used.

The hydrocracking reaction can usually be performed sufficiently by one-stage reaction, but it may also be conducted in a multiple-stage reaction including two or more beds, according to demand.

Also, the hydrocracking reaction may be carried out in the presence of other gas ingredients including inert gases such as gaseous nitrogen, argon, or helium and steam so far as accomplishment of the object

of the present invention is not hindered.

In the manner as described above, the hydrocracking of heavy hydrocarbon oils can be performed efficiently with an increase yield of light hydrocarbon oils such as gas oil, kerosine, and naphtha.

Heavier fractions such as the heavy hydrocarbon oil remaining unreacted and the other remaining ingredients such as hydrogen may optionally be recycled to the reaction system, according to demand.

Deactivated catalyst may undergo regeneration treatment to be used for reaction repeatedly.

The present invention will be described in more detail with reference to the following Examples. These Examples, however, are not to be construed to limit the scope of the invention.

EXAMPLES 1 AND 2 AND COMPARATIVE EXAMPLES 1 TO 4

EXAMPLE 1

(1) Preparation of a Steaming Zeolite

1,000 g of a Y-type zeolite having a Na_2O content of 0.8 % by weight and a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of 5.0 was placed in a rotary kiln and was subjected to self-steaming treatment for 3 hours at 700°C , to obtain a steaming zeolite, referred to as STM zeolite A hereinafter. The properties of the STM zeolite A are shown in Table 1.

(2) Preparation of a Faujasite-type Aluminosilicate

500 g of the STM zeolite A obtained above was suspended in 6,250 cc of pure water and heated to 75°C with stirring, and 4,677 g of a 10 wt% aqueous nitric acid solution was then added into the suspension in 30 minutes. After completion of the addition of nitric acid, the mixture was maintained at 75°C for 30 minutes and was then filtered. The obtained solids were washed with twenty times amount of hot water followed by drying, to obtain the objective faujasite-type aluminosilicate, referred to aluminosilicate A hereinafter. The properties of the aluminosilicate A are shown in Table 1.

(3) Preparation of a Hydrocracking Catalyst for Heavy Hydrocarbon Oils

67 g of the aluminosilicate A obtained above and 189 g of boehmite gel were added into 50 cc of an ion-exchange water, and the resulting mixture was kneaded. The kneaded mixture was adjusted to a water content appropriate for wet-extrusion and was then molded at a molding pressure of 30 kg/cm^2 into granules of 1 mm in diameter and 3 mm in length. The granules were dried at 120°C for 3 hours followed by air-calcination at 500°C for 3 hours, to obtain an aluminosilicate-alumina carrier having an aluminosilicate A content of 65 % by weight.

Subsequently, 45 ml of an aqueous solution containing 13.6 g of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 74.8 g of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ was added to the aluminosilicate-alumina carrier, and the carrier was then vacuum-impregnated with the solution followed by drying at 90°C for 3 hours and calcination at 500°C for 5 hours, to obtain a catalyst of pellet form. The catalyst had a cobalt content of 4 % by weight in CoO and a molybdenum content of 10 % by weight in MoO_3 .

(4) Hydrocracking Reaction of a Heavy Hydrocarbon Oil

After 100 cc of the catalyst containing the aluminosilicate A, which was obtained above, was charged into a high-pressure fixed-bed flow reactor followed by pre-sulfurization according to ordinary method, a Kuwait atmospheric residue [specific gravity: 0.9770 ($15/4^\circ\text{C}$), 343°C fractions (fractions having boiling points of 343°C or higher (the same rule applies correspondingly to the following)): 97 % by volume, the content of sulfur as S: 4.2 % by weight] was passed through the reactor at a reaction temperature of 400°C .

*C, LHSV=0.3 hr⁻¹, a partial pressure of hydrogen of 135 kg/cm², and H₂/oil ratio of 2,000 Nm³/l, to carry out hydrocracking reaction.

The cracking rate of the 343* °C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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EXAMPLE 2

10 An aluminosilicate, referred to as aluminosilicate B hereinafter, was prepared using the STM zeolite A in the same manner as in (2) of Example 1 with the exception that the amount of the 10 wt% aqueous nitric acid solution used was changed to 2,254 g. The properties of the aluminosilicate B are shown in Table 1.

A catalyst containing the aluminosilicate B was prepared in the same manner as in (3) of Example 1 with the exception that the aluminosilicate B was used in place of the aluminosilicate A.

15 Hydrocracking reaction was carried out in the same manner as in (4) of Example 1 with the exception that the catalyst containing the aluminosilicate B was used in place of the catalyst containing the aluminosilicate A.

The cracking rate of the 343* °C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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COMPARATIVE EXAMPLE 1

25 A steaming zeolite, referred to as STM zeolite B hereinafter, was prepared in the same manner as in (1) of Example 1 with the exception that a Y-type zeolite having a SiO₂/Al₂O₃ of 4.6 was used. The properties of the STM zeolite B are shown in Table 1.

An aluminosilicate, referred to as aluminosilicate C hereinafter, was prepared in the same manner as in (2) of Example 1 with the exception that the STM zeolite B was used in place of the STM zeolite A. The properties of the STM zeolite B are shown in Table 1.

30 A catalyst containing the aluminosilicate C was prepared in the same manner as in (3) of Example 1 with the exception that the aluminosilicate C was used in place of the aluminosilicate A.

Hydrocracking reaction was carried out in the same manner as in (4) of Example 1 with the exception that the catalyst containing the aluminosilicate C was used in place of the catalyst containing the aluminosilicate A.

35 The cracking rate of the 343* °C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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COMPARATIVE EXAMPLE 2

An aluminosilicate, referred to as aluminosilicate D hereinafter, was prepared using the STM zeolite A in the same manner as in (2) of Example 1 with the exception that the amount of the 10 wt% aqueous nitric acid solution used was charged to 6,930 g. The properties of the aluminosilicate D are shown in Table 1.

45 A catalyst containing the aluminosilicate D was prepared in the same manner as in (3) of Example 1 with the exception that the aluminosilicate D was used in place of the aluminosilicate A.

Hydrocracking reaction was carried out in the same manner as in (4) of Example 1 with the exception that the catalyst containing the aluminosilicate D was used in place of the catalyst containing the aluminosilicate A.

50 The cracking rate of the 343* °C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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EXAMPLE 3

A steaming zeolite, referred to as STM zeolite C hereinafter, was prepared by charging 1,000 g of a Y-

type zeolite having a Na_2O content of 0.8 % by weight and a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of 5.0 into a rotary kiln followed by self-steaming treatment at 780°C for 3 hours. The properties of the STM zeolite C are shown in Table 1.

5 An aluminosilicate, referred to as aluminosilicate E hereinafter, was prepared in the same manner as in (2) in Example 1 with the exception that the STM zeolite C was used in place of the STM zeolite A. The properties of the aluminosilicate E are shown in Table 1.

A catalyst containing the aluminosilicate E was prepared in the same manner as in (3) of Example 1 with the exception that the aluminosilicate E was used in place of the aluminosilicate A.

10 Hydrocracking reaction was carried out in the same manner as in (4) of Example 1 with the exception that the catalyst containing the aluminosilicate E was used in place of the catalyst containing the aluminosilicate A.

The cracking rate of the 343°C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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COMPARATIVE EXAMPLE 4

20 An aluminosilicate, referred to as aluminosilicate F hereinafter, was prepared using the STM zeolite A in the same manner as in (2) of Example 1 with the exception that the amount of the 10 wt% aqueous nitric acid solution used was changed to 490 g. The properties of the aluminosilicate F are shown in Table 1.

A catalyst containing the aluminosilicate F was prepared in the same manner as in (3) of Example 1 with the exception that the aluminosilicate F was used in place of the aluminosilicate A.

25 Hydrocracking reaction was carried out in the same manner as in (4) of Example 1 with the exception that the catalyst containing the aluminosilicate F was used in place of the catalyst containing the aluminosilicate A.

The cracking rate of the 343°C fractions, the yield of light hydrocarbon oils, the yield of gas fraction, and the desulfurization rate after 400 hours from the beginning of the reaction are shown in Table 2.

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Table 1

		Example Nos.		Comparative Example Nos.			
		1	2	1	2	3	4
5	Properties of parent zeolite (Steaming zeolite)						
10	Lattice constant (Å)	24.39	24.39	24.40	24.39	24.35	24.39
	Bulk SiO ₂ /Al ₂ O ₃ molar ratio	5.0	5.0	4.6	5.0	5.0	5.0
	Framework SiO ₂ /Al ₂ O ₃ molar ratio	17.2	17.2	12.2	17.2	32.0	17.2
	Specific surface area (m ² /g)	599	599	437	599	482	599
15	Absorption percentage of silanol group (3740cm ⁻¹) (%)	3.8	3.8	4.2	3.8	5.6	3.8
	Volume percentage of 8 - 20Å pores (%)	16.3	16.3	14.2	16.3	8.4	16.3
	Crystallinity* (%)	100	100	74	100	73	100
20	Properties of resulting aluminosilicate						
	Lattice constant (Å)	24.30	24.37	24.30	24.21	24.25	24.39
	Crystallinity* (%)	55	90	46	14	84	98
25	Bulk SiO ₂ /Al ₂ O ₃ molar ratio	30.9	8.7	17.3	68.8	49.5	8.2
	Framework SiO ₂ /Al ₂ O ₃ molar ratio	27.7	20.2	18.7	28.6	85.6	17.4
	Specific surface area (m ² /g)	747	752	547	638	586	687
	Absorption percentage A of silanol group (3740cm ⁻¹) (%)	48.4	36.6	13.3	16.3	8.1	5.7
30	Absorption percentage of B of Brønsted acid (3560cm ⁻¹) (%)	19.3	17.4	3.9	2.4	3.3	20.3
	Volume percentage of 8 - 20Å pores (%)	26.3	21.1	20.3	27.5	17.5	17.8
	A/B	2.5	2.1	3.4	6.8	2.5	0.3

* The crystallinity was determined by considering that of STM zeolite A as to be 100 %.

Table 2

		Example Nos.		Comparative Example Nos.			
		1	2	1	2	3	4
45	Percentage of 343 °C fractions cracked (wt%)	58	62	33	35	23	38
	Yield of light hydrocarbon oils * (wt %)	52	55	32	34	23	35
	Yield of gas fractions (wt%)	6	7	1	1	1	3
50	Desulfurizing activity (wt%)	88	87	86	86	85	86

* The light hydrocarbon oils had boiling points of 343 °C or lower.

Claims

1. A novel faujasite-type aluminosilicate which absorbs an infrared in a frequency region of $3740 \pm 10 \text{ cm}^{-1}$ in an absorption percentage A of at least 20 % and absorbs an infrared in frequency region of $3560 \pm 10 \text{ cm}^{-1}$ in an absorption percentage B of at least 5 %, the ratio of A/B being at least 2, has a specific surface area of at least $650 \text{ m}^2/\text{g}$, has a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 20 to 50, and has a lattice constant of from 24.15 to 24.50 \AA .
2. The novel faujasite-type aluminosilicate of claim 1, which absorbs an infrared in the frequency region of $3740 \pm 10 \text{ cm}^{-1}$ in an absorption percentage A of at least 22 % and absorbs an infrared in the frequency region of $3560 \pm 10 \text{ cm}^{-1}$ in an absorption percentage B of at least 7 %, the ratio of A/B being at least 2.1, has a specific surface area of at least $660 \text{ m}^2/\text{g}$, has a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 22 to 48, and has a lattice constant of from 24.20 to 24.50 \AA .
3. A method of producing the faujasite-type aluminosilicate catalyst of claim 1, comprising: treating a faujasite-type zeolite with from 2 to 20 moles of an acid per 1 kg of said faujasite-type zeolite, said faujasite-type zeolite having a framework $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of from 15 to 25, having a specific surface area of at least $500 \text{ m}^2/\text{g}$, and having a lattice constant of from 24.36 to 24.50 \AA .
4. The method of claim 3, wherein said acid is selected from the group consisting of hydrochloric acid, nitric acid, sulfuric acid, and acetic acid.
5. The method of claim 3 or 4, therein said faujasite-type zeolite is treated with said acid at a temperature of from room temperature to 100°C for from 0.1 to 12 hours.
6. The method of claim 3, wherein said faujasite-type zeolite is a steamed faujasite-type zeolite obtainable by treating a faujasite-type zeolite having a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of at least 4.8 and a Na_2O content of not more than 3.0 % by weight with steam at a temperature of from 500 to 900°C for from 0.5 to 5 hours.
7. The method of claim 3, wherein said faujasite-type zeolite is a steamed faujasite-type zeolite obtainable by treating a faujasite-type zeolite having a $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio of at least 4.8 and a Na_2O content of not more than 3.0 % by weight with steam at a temperature of from 500 to 900°C for from 0.5 to 5 hours, said acid is selected from the group consisting of hydrochloric acid, nitric acid, sulfuric acid, and acetic acid, and said faujasite-type zeolite is treated with from 2 to 20 moles of said acid per 1 kg of said faujasite-type zeolite at a temperature of from room temperature to 100°C for from 0.1 to 12 hours.
8. The method of claim 7, wherein said acid is nitric acid and the faujasite-type zeolite to be used for the preparation of the steaming faujasite-type zeolite is a Y-type zeolite.
9. A hydrocracking catalyst for a heavy hydrocarbon oil, comprising: a metallic element belonging to the VIb Group of the Periodic Table and a metallic element belonging to the VIII Group of the Periodic Table which are supported on a carrier containing the faujasite-type aluminosilicate of any of the claims 1 or 2.
10. The hydrocracking catalyst for a heavy hydrocarbon oil of claim 9, wherein said carrier is a composition comprising said faujasite-type aluminosilicate and an inorganic oxide selected from the group consisting of alumina, silica, titania, and alumina-boria, said faujasite-type aluminosilicate being present in the composition in an amount of from 5 to 95 % by weight based on the total of the amounts of said faujasite-type aluminosilicate and the inorganic oxide.
11. The hydrocracking catalyst for a heavy hydrocarbon oil of claim 9 or 10, therein said metallic element belonging to the VIb Group of the Periodic Table is in an amount of from 3 to 24 % by weight based on the total weight of the hydrocracking catalyst and said metallic element belonging to the VIII Group of the Periodic Table is in an amount of from 0.7 to 20 % by weight based on the total weight of the hydrocracking catalyst.
12. The hydrocracking catalyst for a heavy hydrocarbon oil of any of the claims 9 to 11, therein said metallic element belonging to the VIb Group of the Periodic Table is molybdenum or tungsten and said metallic element belonging to the VIII Group of the Periodic Table is cobalt or nickel.
13. The hydrocracking catalyst for a heavy hydrocarbon oil of claim 9, wherein said carrier is a composition comprising said faujasite-type aluminosilicate and an inorganic oxide selected from the group consisting of alumina, silica, titania, and alumina-boria, said faujasite-type aluminosilicate being present in the composition in an amount of from 5 to 95 % by weight based on the total of the amounts of said faujasite-type aluminosilicate and the inorganic oxide, said metallic element belonging to the VIb Group of the Periodic Table is molybdenum or tungsten and is in an amount of from 3 to 24 % by weight based on the total weight of the hydrocracking catalyst, and said metallic element belonging to the VIII Group of the Periodic Table is cobalt or nickel and is in an amount of from 0.7 to 20 % by weight based on the total weight of the hydrocracking catalyst.
14. The hydrocracking catalyst for a heavy hydrocarbon oil of claim 13, wherein the inorganic oxide is alumina, said metallic element belonging to the VIb Group of the Periodic Table is molybdenum, and said metallic element belonging to the VIII Group of the Periodic Table is cobalt.

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